

SCIENTIFIC REPORTS

OPEN

Direct Conversion of $\text{CH}_3\text{NH}_3\text{PbI}_3$ from Electrodeposited PbO for Highly Efficient Planar Perovskite Solar Cells

Received: 23 February 2015

Accepted: 31 July 2015

Published: 29 October 2015

Jin-hua Huang, Ke-jian Jiang, Xue-ping Cui, Qian-qian Zhang, Meng Gao, Mei-ju Su, Lian-ming Yang & Yanlin Song

Organic-inorganic hybrid perovskite materials have recently been identified as a promising light absorber for solar cells. In the efficient solar cells, the perovskite active layer has generally been fabricated by either vapor deposition or two-step sequential deposition process. Herein, electrochemically deposited PbO film is *in situ* converted into $\text{CH}_3\text{NH}_3\text{PbI}_3$ through solid-state reaction with adjacent $\text{CH}_3\text{NH}_3\text{I}$ layer, exhibiting a large-scale flat and uniform thin film with fully substrate coverage. The resultant planar heterojunction photovoltaic device yields a best power conversion efficiency of 14.59% and an average power conversion efficiency of $13.12 \pm 1.08\%$ under standard AM 1.5 conditions. This technique affords a facile and environment-friendly method for the fabrication of the perovskite based solar cells with high reproducibility, paving the way for the practical application.

Thin film solar cells are an important technology to afford cost-competitive solar energy through reduced materials and simplified fabrication process as compared to costly inorganic semiconductor-based photovoltaics¹. Recently, organo-metal hybrid perovskites materials, such as MAPbX_3 ($\text{MA} = \text{CH}_3\text{NH}_3^+$; $\text{X} = \text{Cl}$, Br or I) have attracted tremendous attention for thin film photovoltaics due to their promising features such as solution processability, high crystallinity, direct and tunable band-gaps, and high hole/electron transport ability^{2,3}. In the last five years, the perovskite materials have been successfully employed in both mesoscopic and planar-structured solar cells, exhibiting power conversion efficiencies (PCEs) more than 15%⁴⁻¹². Recently, a certified PCE of 17.9% was reported by the National Renewable Energy Laboratory (NREL)¹³. In addition, a higher efficiency of 19.3% was also achieved through sophisticated interface engineering by Yang's group¹⁴. At this stage, the identification of suitable technology for the production of perovskite solar cells with reduced wastage of toxic Pb material, low cost, and scalability to large area manufacturing would be regarded as the next important milestone.

It was found that the device performance was strongly determined by the morphology and structure of the perovskite active layer, which is in turn relied on the deposition methods¹⁵⁻²⁶. In the early stages, the perovskite materials were deposited on a mesoporous or planar substrate by a single-step spin-coating method using PbX_2 and MAX from a common solvent such as γ -butyrolactone⁹ or dimethylformamide¹⁰. In the process, rapid reaction between the two precursors happened, and resulted in the uncontrolled perovskite precipitation and large morphological variations, leading to a large fluctuation of the device performance. Considering this problem, vapor deposition technique, including dual-source or sequential method, was used to create highly flat and uniform perovskite films, and the resultant devices exhibited high PCEs of more than 15%^{11,24}. This technique, however, required high vacuum and severe control of the operating parameters, hindering large-scale production. On the other hand, two-step approaches

Key Laboratory of Green Printing, Institute of Chemistry, Chinese Academy of Sciences Beijing 100190, P. R. China. Correspondence and requests for materials should be addressed to K.-J.J. (email: kjjiang@iccas.ac.cn) or Y.S. (email: ylsong@iccas.ac.cn)

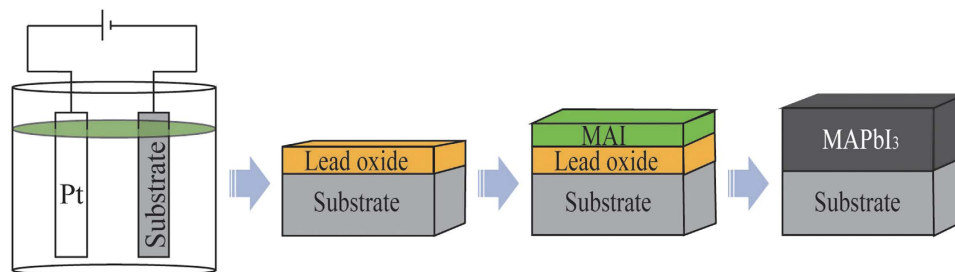


Figure 1. Schematic illustration of the perovskite $\text{CH}_3\text{NH}_3\text{PbI}_3$ film formation on $c\text{-TiO}_2$ coated FTO glass substrate starting from electrodeposited PbO layer and subsequent *in-situ* reaction with adjacent $\text{CH}_3\text{NH}_3\text{I}$ (MAI) layer.

have been successfully investigated to prepare the perovskite film with relatively high surface coverage. The methods include two-step sequential solution process¹⁰, vapor assisted solution method²⁷, and solution-processed precursor stacking method^{28,29}. In the above two-step methods, PbI_2 was first spin-coated on a mesoporous or planar substrate, followed by liquid-solid, gas-solid or solid-solid interdiffusion reaction with MAX. However, it is still a challenge to prepare a large-area, flat and uniform perovskite film using spin-coating technique, which is prerequisite for highly efficient solar cells.

In the previous report, we used an electrochemical method to deposit PbO in mesoscopic TiO_2 films, following with iodination for PbI_2 , and interdiffusion reaction with $\text{CH}_3\text{NH}_3\text{I}$. The as-prepared $\text{CH}_3\text{NH}_3\text{PbI}_3$ was used as a light absorber in the mesoscopic solar cells, exhibiting a high PCE of 12.5%³⁰. In this work, we report that the perovskite can be successfully prepared through direct conversion of the electrodeposited PbO on $c\text{-TiO}_2$ -coated FTO glass substrate by reacting with adjacent $\text{CH}_3\text{NH}_3\text{I}$ layer, allowing for a large-scale flat and uniform thin film with fully substrate coverage. The resultant planar heterojunction photovoltaic device yields a best power conversion efficiency of 14.59% and an average power conversion efficiency of $13.12 \pm 1.08\%$ under standard AM 1.5 conditions.

Electrochemical deposition is a versatile technique for producing surface coatings, possessing precise controllability, low temperature operation, rapid deposition rate, large-scale production capacity, and relatively low cost^{31–36}. Here, we report a facile and efficient one-step method for the fabrication of the perovskite $\text{CH}_3\text{NH}_3\text{PbI}_3$ film starting from electrodeposited PbO for efficient planar solar cells. As shown in Fig. 1, PbO is first electrochemically deposited on a fluorine-doped tin oxide glass substrate coated with a thin compact TiO_2 ($c\text{-TiO}_2$), followed by spin-coating a layer of MAI ($\text{CH}_3\text{NH}_3\text{I}$). The staggered layers were subjected to solid-solid interdiffusion reaction at 150°C for 1 h under N_2 atmosphere, directly *in situ* converting to $\text{CH}_3\text{NH}_3\text{PbI}_3$. Unlike the spin-coating method starting with PbI_2 , the film thickness of electrodeposited PbO can be well controlled, exhibiting a large-area, flat and uniform film, while certain thick PbI_2 film was usually realized by spin coating hot and high concentration of PbI_2 solution on the hot substrate, resulting in an unsmooth film. In addition, the PbO film could be cemented with the $c\text{-TiO}_2$ film, favoring the charge transfer from the converted perovskite to the TiO_2 film. Moreover, a certain amount of PbO can be deposited on the substrate as required, without the waste of toxic Pb-containing materials inevitably generated in the existing preparation techniques^{6,10,11,23}. Using our new method, the resultant device yielded a best PCE of 14.59% under standard AM 1.5 conditions, while a PCE of 12.55% was achieved for the film prepared by the reported solution-processed precursor stacking method²⁸.

Electrodeposition of metal oxide films is usually conducted by cathodic reactions in the aqueous solutions^{31–35}. The formation mechanism of the metal oxide films was postulated as follows: base ion (OH^-) is generated in cathode, then reacts with metal cation, such as Pb^{2+} , to form corresponding metal hydroxide $\text{Pb}(\text{OH})_2$, and finally converted to PbO by dehydration^{31,33–34}. Here, the electrodeposition of PbO film was performed in a single-compartment cell equipped with two electrodes: $c\text{-TiO}_2$ -coated FTO glass as a working electrode, Pt as a counter-electrode. The deposition baths consisted of aqueous solutions of 2 mM $\text{Pb}(\text{CH}_3\text{COO})_2$, 1.5 M dimethyl sulfoxide (DMSO) and 200 mM H_2O_2 . The electrodeposition was carried out at 70°C with a potential of around 1.3 V. The deposited films were rinsed with water and dried in air at room temperature.

The morphology and structure of the perovskite film is crucial for the device performance. Figure 2 presents cross-sectional and top-view images of the electrodeposited PbO and *in situ* converted perovskite on the FTO-glass coated with an 80-nm-thick compact TiO_2 ($c\text{-TiO}_2$) layer. As shown in Fig. 2a, the as-deposited PbO film has smooth surface with full surface coverage over a large area. The high resolution image in Fig. 2b shows that a lot of holes were homogeneously distributed over the whole film with size of ~ 20 nm, which might form from the gas evolution during the deposition process³⁴. The cross sectional image shows that the PbO film is flat over a long range with thickness of ~ 70 nm, as shown in Fig. 2c,d. After the reaction with $\text{CH}_3\text{NH}_3\text{I}$, the perovskite film formed showing smooth surface with crystal grain sizes ranging from 300–800 nm (Fig. 2e,f). The cross sectional images (Fig. 2g,h) show that the perovskite film is flat with thickness of ~ 350 nm, which is 5 times thicker than that of the

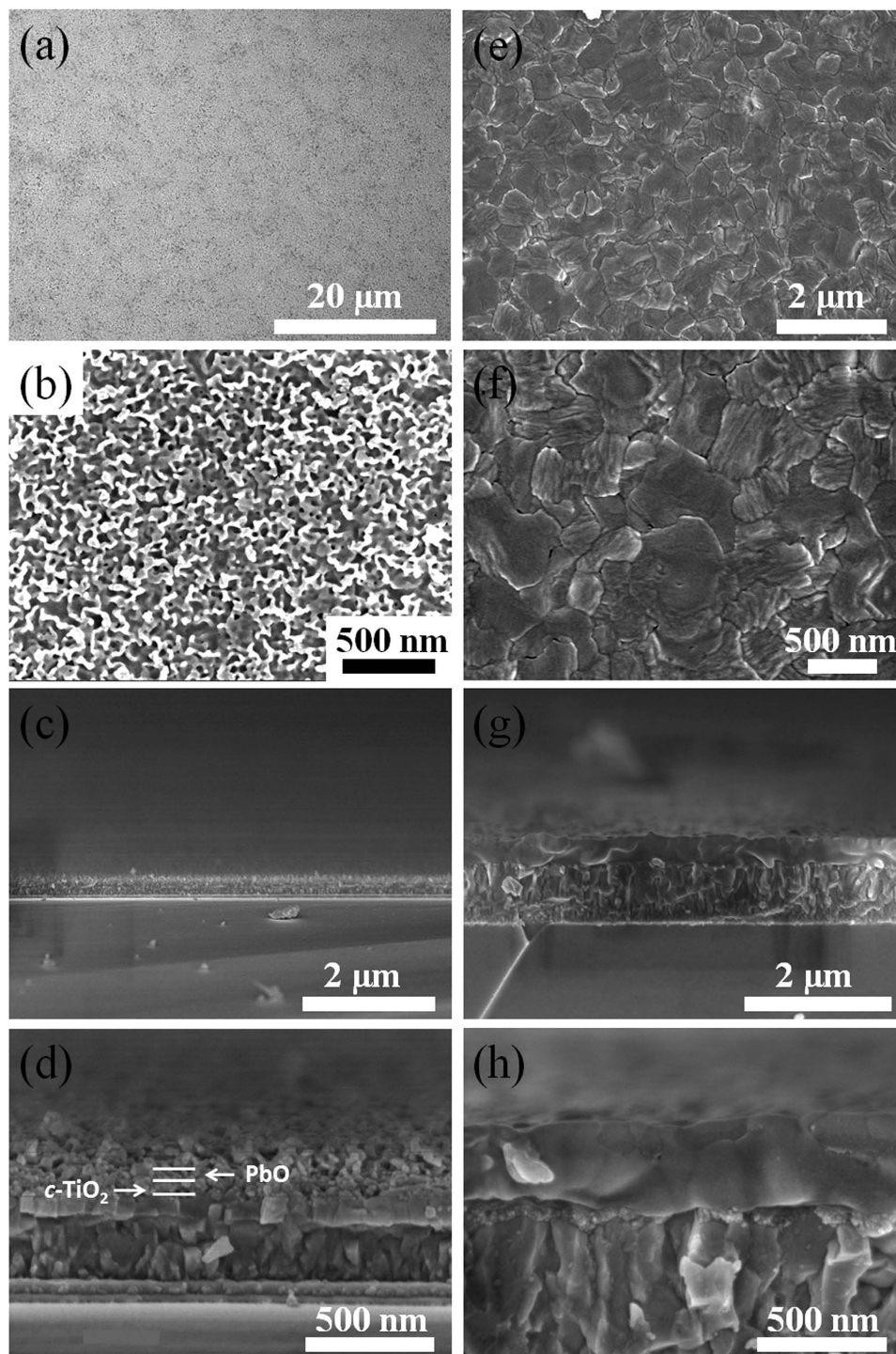


Figure 2. Top-view SEM images of electrodeposited lead oxide film, (a) with low magnification, (b) with high magnification. Cross-sectional SEM images of the lead oxide film on *c*-TiO₂ coated FTO, (c) with low magnification, (d) with high magnification. Top-view SEM images of the perovskite film *in situ* prepared from the reaction of the lead oxide and MAI, (e) with low magnification, (f) with high magnification. Cross-sectional SEM images of the perovskite film on *c*-TiO₂ coated FTO, (g) with low magnification, (h) with high magnification.

corresponding PbO film due to the volume expansion as a result of interdiffusion reaction between the PbO and CH₃NH₃I. According to calculations, the perovskite (density: 4.16 g/cm³) would be about 6.3 times as thick as the lead oxide (density: 9.50 g/cm³) if the latter is completely converted to the former on a flat substrate with the same cross-sectional area. The actual thickness of the perovskite is lower as

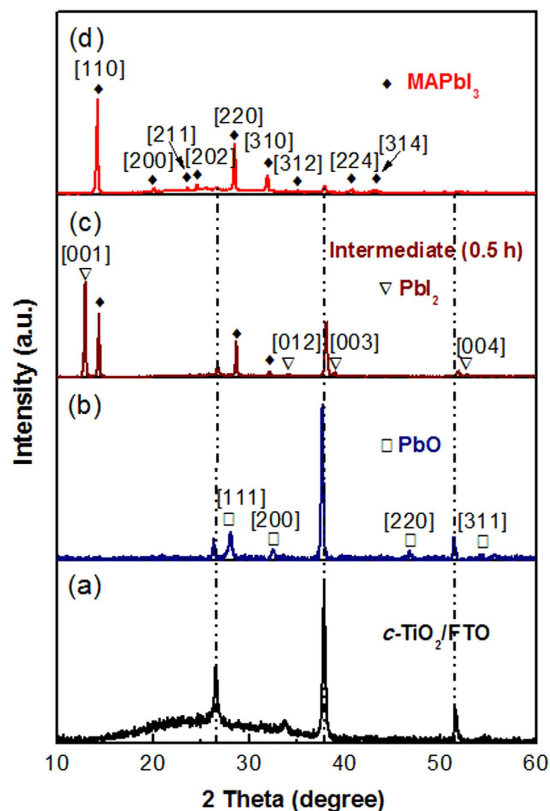


Figure 3. The XRD patterns of (a) the *c*-TiO₂/FTO substrate, (b) the electrodeposited PbO on the substrate, (c) the intermediate on the substrate after 0.5 h interdiffusion reaction, and (d) the perovskite film on the substrate after 1 h interdiffusion reaction.

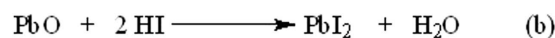
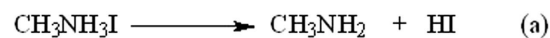


Figure 4. The mechanism to form the perovskite from PbO and CH₃NH₃I.

compared to the calculated value, which is most likely due to the presence of the holes in the PbO film (Fig. 2b).

X-ray diffraction (XRD) was used for the phase identification of the electrodeposited PbO and *in situ* converted perovskite films on the *c*-TiO₂/FTO substrate. XRD pattern of the *c*-TiO₂/FTO substrate was recorded for comparison, as shown in Fig. 3a. After the electrodeposition, new peaks appeared at 28.2°, 32.7°, 46.9°, and 55.7°, which are assigned to [111], [200], [220], and [311] of lead oxide crystal planes, corresponding to the data in JCPDS card no. 27–1201. After coating CH₃NH₃I on the PbO surface and heating at 150 °C for 30 minutes, the peaks for the PbO disappeared, and new peaks were observed. The peaks located at 12.6°, 34.3°, 39.5°, and 52.3° are assigned to (001), (012), (003), and (004) lattice planes of the 2H polytype PbI₂ (JCPDS card No. 73–1750)¹⁰, while the peaks at 14.2°, 20.0°, 23.6°, 24.6°, 28.5°, 31.9°, 35.0°, 40.6°, and 43.2°, are corresponding to the reflections from [110], [200], [211], [202], [220], [310], [312], [224], and [314] of the perovskite²⁷. After the reaction lasted for 1 h, the peaks for PbI₂ disappeared, and the intensity of the peaks from the perovskite became stronger without other new peaks, indicating the complete transformation from the electrodeposited PbO to the perovskite CH₃NH₃PbI₃ through the *in situ* solid-solid reaction. The phase evolution indicates the formation mechanism of the perovskite as follows:

The CH₃NH₃I gradually decomposed to CH₃NH₂ and HI at the elevated temperature at the initial stage (Fig. 4a), and the generated HI further reacted with the PbO to form PbI₂ (Fig. 4b). Finally the PbI₂ reacted with CH₃NH₃I to form the perovskite CH₃NH₃PbI₃ (Fig. 4c), similar to the Huang's report²⁷. In the processes, PbI₂ is an intermediate. It should be noted that equivalent amount of H₂O was generated

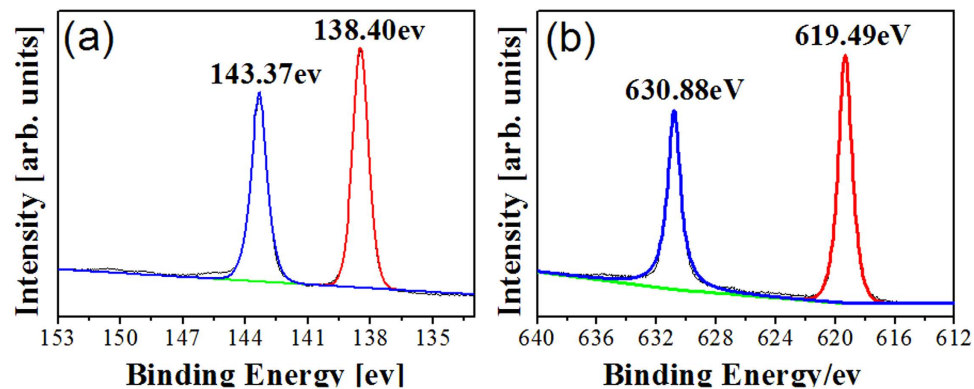


Figure 5. Pb 4f (a) and I 3d (b) core level spectra of the as-prepared perovskite surface measured with a photonenergy of 4000 eV.

during the conversion process from PbO to PbI₂ (Eq-b), and the water could have positive effect on the formation of the provskite layer and thus improve the photovoltaic performance^{37,38}.

XPS was used to determine the chemical environment and stoichiometry at the surface of the perovskite film. Pb 4f and I 3d core level spectra were shown in Fig. 5, and the whole spectra were presented in Fig. S1. In the case of Pb 4f, the binding energies of Pb 4f_{7/2} and Pb 4f_{5/2} were observed at 138.40 eV and 143.37 eV with a spin-orbit split of 4.97 eV. For I 3d, the energies of I 3d_{5/2} and I 3d_{3/2} were found at 619.49 eV and 630.88 eV with a spin-orbit split of 11.39 eV. All the data are well consistent with those reported previously for CH₃NH₃PbI₃ film³⁶. In addition, the stoichiometry between lead and iodine was roughly assessed to be about 1:3 through the integrated intensity of the peaks for Pb 4f and I 3d core levels. For further investigation of the evolution, EDX mapscan was performed for the cross sectional perovskite film, as shown in Fig. S2. The result shows that the atomic ratio for lead and iodine is about 1:3 for the CH₃NH₃PbI₃ film. In combination with the XRD results, it is believed that the electrodeposited PbO film could be completely *in situ* converted to the perovskite through solid-solid interdiffusion reaction with CH₃NH₃I. Recently, PbO coated TiO₂ films were fabricated by calcination of the film prepared from a mixture of TiO₂ colloid and PbI₂ or Pb(CH₃COO)₂. The as-prepared PbO on the films were further transferred to PbI₂ through reaction with HI. The resultant PbI₂-coated TiO₂ films were used for the preparation of the perovskite-decorated TiO₂ films. In the reports, it was observed that the perovskite had poor coverage on TiO₂ particles or nanofibers in the films^{39,40}.

In our experiments, the resultant perovskite film is flat and uniform with full surface coverage, which is prerequisite for high efficient perovskite solar cells. We also tried to prepare PbI₂ film by spin coating method, and then *in situ* being converted to the perovskite according to the reported method²⁷. We found it difficult to get a flat and uniform perovskite film (as shown in Fig. S3). In addition, other soluble Pb-salts such as Pb(CH₃COO)₂, Pb(NO₃)₂, Pb(ClO₄)₂ can be used for the electrodeposition of PbO film³¹⁻³⁴. In the electrodeposition, they can be completely converted to the PbO without waste. Thus, the present method is environment-friendly for the fabrication of the perovskite active film for the photovoltaic devices.

The resultant perovskite film was used as light harvester for fabrication of planar heterojunction solar cells. The device (denoted as **device 1**) was constructed with a structure of FTO/*c*-TiO₂ (~80 nm)/CH₃NH₃PbI₃ (~350 nm)/Spiro-OMeTAD (150 nm)/Au (80 nm) (see Fig. S4). For comparison, another device was prepared with the same architecture except for the perovskite active layer fabricated starting from PbI₂ according to the previous report (denoted as **device 2**)²⁸. Both the devices were characterized by recording photocurrent density-voltage (*J*-*V*) curves with reverse (from *V*_{oc} to *J*_{sc}) or forward (from *J*_{sc} to *V*_{oc}) bias scanning at a rate of 100 mV/s under AM 1.5 irradiation (100 mW cm⁻²), and the photovoltaic parameters were listed in Table 1. As shown in Fig. 6a, the best performing **device 1** gave a short circuit photocurrent density (*J*_{sc}) of 20.97 mA/cm², an open circuit voltage (*V*_{oc}) of 0.98 V, a fill factor (*FF*) of 0.71, and a power conversion efficiency (PCE) of 14.59% with the reversed scanning, and the average PCE of **device 1** is 13.12% for 16 samples with a relatively low standard deviation of 8% (shown in Fig. S6a and Table S1). In the case of the reverse bias scanning, slight hysteresis effect was observed for the **device 1**, where the best performing **device 1** gave a relatively lower PCE of 13.90% due to the decreased *FF* (0.69 vs. 0.71). With the reverse bias scanning, the best performing **device 2** gave a *J*_{sc} of 19.81 mA/cm², a *V*_{oc} of 0.96 V, a *FF* of 0.66, and a PCE of 12.55%, with an average PCE of 9.96% and a standard deviation of 15%, as shown in Fig. S5, S6b and Table S2. In addition, obvious hysteresis effect was observed for the **device 2**, where the best performing **device 2** gave a decreased PCE of 11.2% under the forward scanning. The results indicate the **device 1** has a higher PCE and reproducibility as compared to the **device 2**, which can be explained due to the high-quality perovskite film evolved from the flat and uniform electrodeposited PbO film. Figure 6b shows the external quantum efficiency (*EQE*) spectrum of **device 1**. A broad spectral response in the range of 300–820 nm was observed with the

Device		$J_{sc}/\text{mA cm}^{-2}$	V_{oc}/V	FF	PCE/%
1	Rev. ^[a]	20.97	0.98	0.71	14.59
	For. ^[b]	20.98	0.96	0.69	13.90
	av. ^[c]	19.81 ± 1.57	0.96 ± 0.03	0.69 ± 0.02	13.12 ± 1.08
2	Rev. ^[a]	19.81	0.96	0.66	12.55
	For. ^[b]	19.85	0.91	0.62	11.20
	av. ^[c]	17.82 ± 2.18	0.90 ± 0.05	0.62 ± 0.13	9.96 ± 1.52

Table 1. Photovoltaic performance of device 1 and device 2 measured with reverse and forward bias scanning at a rate of 100 mV/s under AM 1.5 irradiation (100 mW cm^{-2}) under the 100 mW cm^{-2} AM1.5G illumination. ^[a]reverse scan; ^[b]forward scan; ^[c]average and standard deviation for a batch of 16 samples.

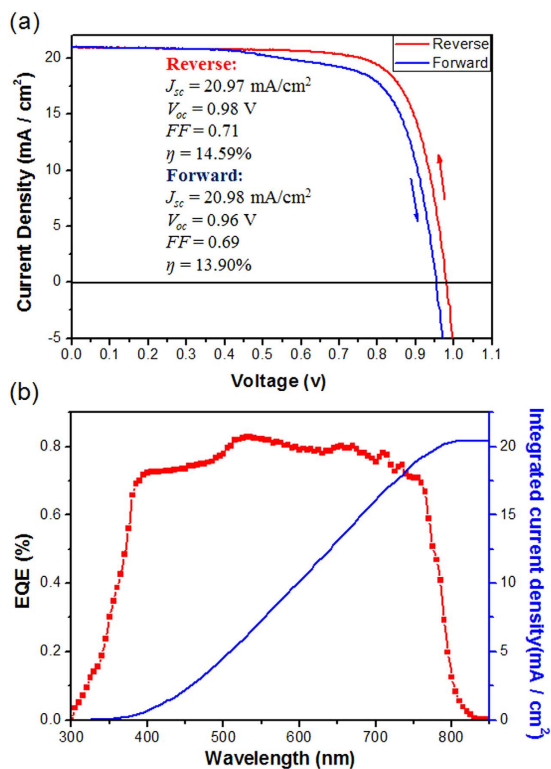


Figure 6. (a) J - V characteristics of device 1 with reverse (from V_{oc} to J_{sc}) or forward (from J_{sc} to V_{oc}) bias scanning at a rate of 100 mV/s under AM 1.5 irradiation (100 mW cm^{-2}) under AM 1.5G illumination; (b) the EQE spectrum (red) and the integrated photocurrent density (blue) of device 1 expected to be generated under AM 1.5G irradiation.

highest value of 82% at 530 nm. The integrated current density (20.44 mA/cm^{-2}) is well consistent with the measured value of the J_{sc} .

In summary, we developed a facile and environment-friendly technology for the fabrication of long-range continuous and homogeneous perovskite $\text{CH}_3\text{NH}_3\text{PbI}_3$ film with full surface coverage. In the method, electrodeposited PbO films can be *in situ* directly converted to the perovskite through solid-solid interdiffusion reaction with $\text{CH}_3\text{NH}_3\text{I}$. We proposed a plausible mechanism for the formation of uniform and dense provskite layer. The resulting perovskite film was used as light harvester to construct planar heterojunction solar cells, giving an initial power conversion efficiency of 14.59%, which is among the highest values. Further efficiency enhancement can be expected following the optimization of the film thickness, morphology and relative interface engineering. The combination of simple control of deposit thickness, low processing temperature, low cost of equipment, large-scale, high reproducibility and environmental friendly process, will make this method as a promising technique for the practical production of the perovskite and other thin film photovoltaics.

Methods

Substrate preparation was carried out under ambient conditions. FTO-coated glass (14 ohm/sq) was patterned by etching with Zn powder and 2 M HCl diluted in deionized water. The substrates were then cleaned with detergent diluted in deionized water, rinsed with deionized water, acetone and ethanol, and dried with clean dry air. After oxygen plasma treatment, the clean substrates were spin-coated with 0.15 M and 0.3 M titanium diisopropoxide bis(acetylacetonate) at 3,000 r.p.m. for 30 s subsequently. After drying at 125 °C for 10 min, they were sintered at 500 °C for 60 min in air. The substrate was immersed in 40 mM TiCl₄ aqueous solutions at 70 °C for 30 min and washed with distilled water and ethanol, followed by annealing at 500 °C for 30 min in air to form a compact layer of TiO₂ (*c*-TiO₂).

The lead oxide film was prepared according to the following method: Electrodeposition of thin films was performed in a single-compartment cell equipped with two electrodes: FTO/*c*-TiO₂ glass as a working electrode, Pt as a counter-electrode. The deposition baths consisted of aqueous solutions of 2 mM Pb(CH₃COO)₂, 1.5 M DMSO and 200 mM H₂O₂ (It needs to bubble sufficient nitrogen gas to displace the dissolved air before adding H₂O₂). The lead oxide film was carried out by potentiostatic electrolysis, the thickness and quality of the film is optimized by controlling the deposition parameters such as time, deposition potential (1.3 V), the deposition temperature (70 °C), etc. The deposited films were rinsed with water and dried under air at room temperature. After drying, 30 μL MAI solutions (50 mg/mL) were spun on the top of the film at 4,000 rpm for 30 s. The films were placed in a petridish and annealed at 150 °C for 1 h. After cooling down, the isopropanol was spun on the as-prepared films at 4,000 rpm for 30 s and dried at 70 °C. then, the spiro-OMeTAD-based hole-transfer layer (80 mg spiro-OMeTAD, 28.5 μL 4-tertbutylpyridine and 17.5 μL lithium-bis(trifluoromethanesulfonyl)-imide (Li-TFSI) solution (520 mg Li-TFSI in 1 ml acetonitrile) all dissolved in 1 ml chlorobenzene) was deposited by spin coating at 4,000 r.p.m. for 30 s. Finally, the counter Au electrode was deposited by thermal evaporation at a base pressure of 5 × 10⁻⁵ mbar. The active area was 0.04 cm².

References

- Chopra, K. L. *et al.* Thin-film solar cells: an overview. *Prog. Photovolt: Res. Appl.* **12**, 69–92 (2004).
- Stranks, S. D. *et al.* Electron-hole diffusion lengths exceeding 1 micrometer in an organometal trihalide perovskite absorber. *Science* **342**, 341–344 (2013).
- Xing, G. *et al.* Long-range balanced electron and hole-transport lengths in organic-inorganic CH₃NH₃PbI₃. *Science* **342**, 344–347 (2013).
- Kojima, A. *et al.* Organometal halide perovskites as visible-light sensitizers for photovoltaic cells. *J. Am. Chem. Soc.* **131**, 6050–6051 (2009).
- Lee, M. M. *et al.* Efficient hybrid solar cells based on meso-structured organometal halide perovskites. *Science* **338**, 643–647 (2012).
- Kim, H.-S. *et al.* Lead iodide perovskite sensitized all-solid-state submicron thin film mesoscopic solar cell with efficiency exceeding 9%. *Sci. Rep.* **2**, 591–597 (2012).
- Etgar, L. *et al.* Mesoscopic CH₃NH₃PbI₃/TiO₂ heterojunction solar cells. *J. Am. Chem. Soc.* **134**, 17396–17399 (2012).
- Noh, J. H. *et al.* Chemical management for colorful, efficient, and stable inorganic–organic hybrid nanostructured solar cells. *Nano Lett.* **13**, 1764–1769 (2013).
- Heo, J. H. *et al.* Efficient inorganic–organic hybrid heterojunction solar cells containing perovskite compound and polymeric hole conductors. *Nat. Photon.* **7**, 486–491 (2013).
- Burschka, J. *et al.* Sequential deposition as a route to high-performance perovskite-sensitized solar cells. *Nature* **499**, 316–319 (2013).
- Liu, M. *et al.* Efficient planar heterojunction perovskite solar cells by vapour deposition. *Nature* **501**, 395–398 (2013).
- Liu, D. *et al.* Perovskite solar cells with a planar heterojunction structure prepared using room-temperature solution processing techniques. *Nat. Photon.* **8**, 133–138 (2014).
- NREL, (http://www.nrel.gov/ncpv/images/efficiency_chart.jpg), accessed on Oct. 15, 2014.
- Zhou, H. *et al.* Interface engineering of highly efficient perovskite solar cells. *Science* **345**, 542–546 (2014).
- Conings, B. *et al.* Perovskite-based hybrid solar cells exceeding 10% efficiency with high reproducibility using a thin film sandwich approach. *Adv. Mater.* **26**, 2041–2046 (2014).
- Dualeh, A. *et al.* Effect of annealing temperature on film morphology of organic–inorganic hybrid perovskite solid-state solar cells. *Adv. Funct. Mater.* **24**, 3250–3258 (2014).
- Liang, P.-W. *et al.* Additive enhanced crystallization of solution-processed perovskite for highly efficient planar-heterojunction solar cells. *Adv. Mater.* **26**, 3748–3754 (2014).
- Jeon, N. J. *et al.* Solvent engineering for high-performance inorganic–organic hybrid perovskite solar cells. *Nat. Mater.* **13**, 897–903 (2014).
- Zhao, Y. *et al.* CH₃NH₃Cl-assisted one-step solution growth of CH₃NH₃PbI₃: structure, charge-carrier dynamics, and photovoltaic properties of perovskite solar cells. *J. Phys. Chem. C* **118**, 9412–9418 (2014).
- Wu, Y. *et al.* Retarding the crystallization of PbI₂ for highly reproducible planar-structured perovskite solar cells via sequential deposition. *Energy Environ. Sci.* **7**, 2934–2938 (2014).
- Zuo, C. *et al.* An 80.11% FF record achieved for perovskite solar cells by using the NH₄Cl additive. *Nanoscale*, **6**, 9935–9938 (2014).
- Yu, H. *et al.* The role of chlorine in the formation process of “CH₃NH₃PbI_{3-x}Cl_x” perovskite. *Adv. Funct. Mater.* **24**, 7102–7108 (2014).
- Docampo, P. *et al.* Solution deposition-conversion for planar heterojunction mixed halide perovskite solar cells. *Adv. Energy Mater.* **4**, 1400355 (2014).
- Barrows, A. T. *et al.* Efficient planar heterojunction mixed-halide perovskite solar cells deposited via spray deposition. *Energy Environ. Sci.* **7**, 2944–2950 (2014).
- Chen, C.-W. *et al.* Efficient and uniform planar-type perovskite solar cells by simple sequential vacuum deposition. *Adv. Mater.* **26**, 6647–6652 (2014).
- Xiao, M. *et al.* A fast deposition-crystallization procedure for highly efficient lead iodide perovskite thin-film solar cells. *Angew. Chem. Int. Ed.* **53**, 9898–9903 (2014).

27. Chen, Q. *et al.* Planar heterojunction perovskite solar cells via vapor-assisted solution process. *J. Am. Chem. Soc.* **136**, 622–625 (2014).
28. Xiao, Z. *et al.* Efficient, high yield perovskite photovoltaic devices grown by interdiffusion of solution-processed precursor stacking layers. *Energy Environ. Sci.* **7**, 2619–2623 (2014).
29. Xiao, Z. *et al.* Solvent annealing of perovskite-induced crystal growth for photovoltaic-device efficiency enhancement. *Adv. Mater.* **26**, 6503–6509 (2014).
30. Cui, X.-P. *et al.* Electrodeposition of PbO and its *in situ* conversion to CH₃NH₃PbI₃ for mesoscopic perovskite solar cells. *Chem. Commun.* **51**, 1457–1460 (2015).
31. Zhitomirsky, I. *et al.* Electrochemical preparation of PbO films. *J. Mater. Sci. Lett.* **14**, 807–810 (1995).
32. Sawatani, S. *et al.* Formation of highly crystallized β-PbO thin film by cathodic electrodeposition of Pb and its rapid oxidation in air. *Adv. Funct. Mater.* **15**, 297–302 (2005).
33. Dwivedi, V. K. *et al.* Direct deposition of inorganicorganic hybrid semiconductors and their template-assisted microstructures. *Mater. Chem. Phys.* **137**, 941–946 (2013).
34. Peulon, S. *et al.* Cathodic electrodeposition from aqueous solution of dense or open-structured zinc oxide films. *Adv. Mater.* **8**, 166–170 (1996).
35. Qian, J. *et al.* A selenium-based cathode for a high-voltage tandem photoelectrochemical solar cell *Angew. Chem. Int. Ed.* **51**, 10351–10354 (2012).
36. Lindblad, R. *et al.* Electronic structure of TiO₂/CH₃NH₃PbI₃ Perovskite solar cell interfaces. *J. Phys. Chem. Lett.* **5**, 648–653 (2014).
37. You, J. *et al.* Moisture assisted perovskite film growth for high performance solar cells. *Appl. Phys. Lett.* **105**, 183902–183905 (2014).
38. Ko, H.-S. *et al.* 15.76% efficiency perovskite solar cells prepared under high relative humidity: importance of PbI₂ morphology in two-step deposition of CH₃NH₃PbI₃. *J. Mater. Chem. A* **3**, 8808–8815 (2015).
39. Xiao, Y. *et al.* Preparation of high performance perovskitesensitized nanoporous titanium dioxide photoanodes by *in situ* method for use in perovskite solar cells. *J. Mater. Chem. A* **2**, 16531–16537 (2014).
40. Xiao, Y. *et al.* Electrospun lead-doped titanium dioxide nanofibers and the *in situ* preparation of perovskite-sensitized photoanodes for use in high performance perovskite solar cells. *J. Mater. Chem. A* **2**, 16856–16862 (2014).

Acknowledgements

The authors greatly appreciate the financial support from the National Natural Science Foundation of China (Grant Nos. 61405207, 21174149, 21102150, 51173190, 21073203, 21076002 and 21121001), the National 863 Program (No. 2011AA050521), the 973 Program (2009CB930404, 2011CB932303, 2011CB808400), and the Scientific Equipment Program, ACS (YZ201106).

Author Contributions

J.-H.H., K.-J.J. and Y.L.S. developed the basic concept. J.-H.H., K.-J.J., L.-M.Y. and Y.L.S. contributed to the design of experiment, analysis of the data, and writing of the manuscript. J.-H.H. carried out the experiments for device fabrication, optimization and devices performance measurements. X.-P.C. carried out synthesis of CH₃NH₃I. M.G., Q.-Q.Z. and M.-J.S. carried out XPS and SEM-EDX measurements. All the authors discussed the results and commented on the manuscript.

Additional Information

Supplementary information accompanies this paper at <http://www.nature.com/srep>

Competing financial interests: The authors declare no competing financial interests.

How to cite this article: Huang, J.-H. *et al.* Direct Conversion of CH₃NH₃PbI₃ from Electrodeposited PbO for Highly Efficient Planar Perovskite Solar Cells. *Sci. Rep.* **5**, 15889; doi: 10.1038/srep15889 (2015).



This work is licensed under a Creative Commons Attribution 4.0 International License. The images or other third party material in this article are included in the article's Creative Commons license, unless indicated otherwise in the credit line; if the material is not included under the Creative Commons license, users will need to obtain permission from the license holder to reproduce the material. To view a copy of this license, visit <http://creativecommons.org/licenses/by/4.0/>